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reaction using symbols and formulae of the substances is called Chemical Equation. Example: $A + B \rightarrow C + D$ In this equation, A and B are called reactants and C and D are called the products. The arrow shows the direction of the chemical reaction.

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The chapter of class 10 on chemical reactions and equations is therefore essential for helping students in forming a strong foundational base for higher classes. It explains in detail the elementary concepts of a chemical reaction and chemical equation. Students will learn about balancing a chemical equation after going through the solutions.

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EXOTHERMIC REACTIONS • An exothermic reaction is a chemical reaction that releases energy by light or heat. It is the opposite of an endothermic reaction. • Expressed in a chemical equation: reactants → products + energy For example
1) $4\text{Fe(s)} + 3\text{O}_2\text{(g)} \rightarrow 2\text{Fe}_2\text{O}_3\text{(s)}$ 2) Making ice cubes 3) Mixing sodium sulfite and 30.

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Chemical Equation: A mathematical statement which represents the chemical reaction. Types of Chemical Reactions: Combustion reaction – The reaction of a combustible material with an oxidiser to give an oxidised product. Decomposition reaction – The reaction in which a single component breaks down into multiple

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products.

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Balanced chemical equation The chemical equation that shows the chemical reaction needs to be balanced. A balanced chemical equation occurs when the number of the atoms involved in the reactants side is equal to the number of atoms in the products side. $Zn + H_2SO_4 \rightarrow ZnSO_4 + H_2$ $3Fe (s) + 4H_2O (g) \rightarrow Fe_3O_4 (s) + 4H_2 (g)$

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