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inversely proportionality to the square root of their densities.
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electrolytes or solutes in solution. • the partial vapor pressure of the solvent of an ideal solution is equal to the product of the mole fraction of the solvent and its vapor pressure in pure state.

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Dr Valizadeh (Teaching Materials) - Pharmacy faculty
Objectives of the lecture After completion of this lecture, the student should be able to: 1. Understand the various types of

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pharmaceutical solutions. 2. Define solubility, saturated & unsaturated solutions and polar & non polar solvents. 3. Describe complete & partial miscibility. 4.

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Physical pharmacy. 1. PHYSICAL PHARMACY. 2. Matter: It's what the world is made of. 3. Kinetic Theory of Matter Matter is made up of particles which are in continual random motion. 4. Binding Forces between Molecules The Three States of Matter: In order for molecules

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to exist in aggregates in gases, liquids and solids
Intermolecular forces must exist 4.

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- Physical properties - a characteristic that can be observed

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or measured without changing the identity or composition of the substance • Physical properties used to describe matter can be classified as: 1) Extensive – depends on the . amount. of matter in the sample - e.g. Mass, volume, length . 2) Intensive – depends on the . type. of

THE PROPERTIES AND STRUCTURE OF MATTER

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